

**Useful Stuff**

$$1 \text{ Joule} = 1 \text{ Kg}\cdot\text{m}^2/\text{s}^2$$

$$R = 0.08206 \text{ L}\cdot\text{Atm}/\text{K}\cdot\text{Mole}$$

$$760 \text{ torr} = 1 \text{ atm}$$

$$\text{Avogadro's \#} = 6.022 \times 10^{23}$$

Bond Energy Table and  
Periodic Table at Back of Exam